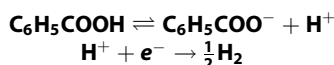


Measuring the solubility of benzoic acid in room temperature ionic liquids using chronoamperometric techniques

Weisi He^a, Debbie S. Silvester^a, Ian Streeter^a, Leigh Aldous^b, Christopher Hardacre^b and Richard G. Compton^{a*}

The electrochemical reduction of benzoic acid (BZA) has been studied at platinum micro-electrodes (10 and 2 μm diameters) in acetonitrile (MeCN) and six room temperature ionic liquids (RTILs): [C₂mim][NTf₂], [C₄mim][NTf₂], [C₄mpyrr][NTf₂], [C₄mim][BF₄], [C₄mim][NO₃] and [C₄mim][PF₆] (where [C_nmim]⁺ = 1-alkyl-3-methylimidazolium, [NTf₂]⁻ = bis(trifluoromethylsulphonyl)imide, [C₄mpyrr]⁺ = *N*-butyl-*N*-methylpyrrolidinium, [BF₄]⁻ = tetrafluoroborate, [NO₃]⁻ = nitrate and [PF₆]⁻ = hexafluorophosphate). Based on the theoretical fitting to experimental chronoamperometric transients in [C₄mpyrr][NTf₂] and MeCN at several concentrations and on different size electrodes, it is suggested that a fast chemical step precedes the electron transfer step in a CE mechanism (given below) in both RTILs and MeCN, leading to the appearance of a simple one-electron transfer mechanism.



The six RTIL solvents and MeCN were saturated with BZA, and potential-step chronoamperometry revealed diffusion coefficients of 170, 4.6, 3.2, 2.7, 1.8, 0.26 and 0.96 × 10⁻¹¹ m² s⁻¹ and solubilities of 850, 75, 78, 74, 220, 2850 and 48 mM in MeCN and the six ionic liquids, respectively, at 298 K. The high solubility of BZA in [C₄mim][NO₃] may suggest a strong interaction of the dissolved proton with the nitrate anion. Although there are relatively few literature reports of solubilities of organic solutes in RTILs at present, these results suggest the need for further studies on the solubilities of organic species (particularly acids) in RTILs, because of the contrasting interaction of dissolved species with the RTIL ions. Chronoamperometry is suggested as a convenient methodology for this purpose. Copyright © 2008 John Wiley & Sons, Ltd.

Keywords: benzoic acid; electrochemical reduction; cyclic voltammetry; chronoamperometry; solubility; diffusion coefficients; room temperature ionic liquids; acetonitrile

INTRODUCTION

Room temperature ionic liquids (RTILs) are compounds composed entirely of ions and exist in the liquid state at and around 298 K.^[1-4] RTILs have several unique properties such as low volatility (considered 'greener' than volatile organic solvents),^[1,5,6] high thermal stability and can be 'tunable' by simply varying the nature of the cation and anion.^[7-9] RTILs have a wide electrochemical window which allows experiments to be performed that are normally out of the range of traditional solvents for example electrodeposition of metals and semiconductors.^[10] They have a high intrinsic conductivity (composed entirely of charge carriers), allowing for electrochemical experiments to be performed without added supporting electrolyte, simplifying experimental set-up and minimizing waste. They are currently being used in other applications such as organic and inorganic synthesis,^[11-13] industrial applications,^[6,14] catalysis,^[11] sensors,^[15] solar cells,^[16] fuel cells^[17] and capacitors.^[18,19]

The exact nature and relative strength of acidic protons dissolved in RTIL systems is still a topic of investigation. For example, conventional acid-base indicators have been used to

physically probe the inherent acidity of ionic liquids, but the observed colour changes are the result of a combination of hydrogen-bonding interactions and the acid-base equilibria.^[20-22] Hardacre *et al.*^[23] demonstrate that Friedel-Crafts reactions in ionic liquids based upon [NTf₂]⁻ (bis(trifluoromethylsulphonyl)imide), the acid catalysts H[BF₄] and H[OTf] (where [BF₄]⁻ = tetrafluoroborate and [OTf]⁻ = trifluoromethylsulphonate) demonstrate significantly lower reaction rates than H[NTf₂], despite all three being strong acids in water. Del Pópolo *et al.*^[24] simulated the dissolution and solvation of HCl in the ionic liquid

* Correspondence to: R. G. Compton, Department of Chemistry, Physical and Theoretical Chemistry Laboratory, Oxford University, South Parks Road, Oxford, OX1 3QZ, UK.
E-mail: richard.compton@chem.ox.ac.uk

a W. He, D. S. Silvester, I. Streeter, R. G. Compton
Department of Chemistry, Physical and Theoretical Chemistry Laboratory, Oxford University, South Parks Road, Oxford, OX1 3QZ, UK

b L. Aldous, C. Hardacre
School of Chemistry and Chemical Engineering/QUILL, Queen's University Belfast, Belfast, Northern Ireland, BT9 5AG, UK

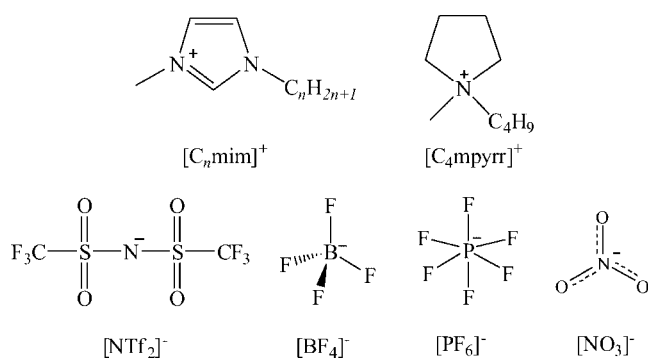


Figure 1. Structures of all the anions and cations employed in the RTIL solvents in this study

[C₁mim]Cl, and observed the rapid formation of the anion [HCl₂][−], rather than dissociation. Several other works have focussed on the behaviour of Brønsted acids in RTILs.^[25–27] Recently, we have reported the oxidation of hydrogen gas^[28,29] and ammonia gas^[30] in RTILs and noted voltammetric behaviour which was mainly dependent on the nature of the RTIL anion. In particular, no reduction waves were observed following the oxidation of H₂ in [C₆mim]Cl and [C₄mim][NO₃],^[29] suggesting the possible formation of [HCl₂][−] and H[NO₃]₂[−].

In this work, the reduction of benzoic acid (BZA) is reported in acetonitrile (MeCN) and in six commonly used RTILs: [C₂mim][NTf₂], [C₄mim][NTf₂], [C₄mpyrr][NTf₂], [C₄mim][BF₄], [C₄mim][NO₃] and [C₄mim][PF₆] (the structures of which are given in Fig. 1) in order to further try to understand the nature of acid/base behaviour in RTIL media. The electrochemical reduction of weak acids is thought to proceed by the following generic CE mechanism:^[31–34]



Cyclic voltammetry for BZA reduction has previously been reported in water/acidic media,^[33–35] and in the aprotic solvents dimethylformamide (DMF)^[31] and dimethylsulphoxide (DMSO),^[32] appearing to be relatively broad and electrochemically quasi-reversible/irreversible in all cases, and to follow the CE reaction mechanism. The kinetics of the homogeneous step (Eqn (1)) can be studied by observing how the limiting current responds to changes in electrode size and concentration of BZA.^[36–39] In this work, the findings from this type of kinetic analysis are discussed, and the results may have implications both in the understanding of electrode kinetics of CE reactions, and generally in the behaviour of protons in RTILs.

EXPERIMENTAL

Chemical reagents

1-Ethyl-3-methylimidazolium bis(trifluoromethylsulphonyl)imide ([C₂mim][NTf₂]), 1-butyl-3-methylimidazolium bis(trifluoromethylsulphonyl)imide ([C₄mim][NTf₂]), *N*-butyl-*N*-methylpyrrolidinium bis(trifluoromethylsulphonyl)imide ([C₄mpyrr][NTf₂]) and their bromide salt precursors were prepared by standard literature procedures.^[40,41] 1-Butyl-3-methylimidazolium tetrafluoroborate

([C₄mim][BF₄] high purity) and 1-butyl-3-methylimidazolium hexafluorophosphate ([C₄mim][PF₆] high purity) were kindly donated by Merck KGaA, and were used as received. The synthesis of 1-butyl-3-methylimidazolium nitrate ([C₄mim][NO₃]) was adapted from a previously published procedure.^[42] AgNO₃ (5.36 g, 0.032 M) and [C₄mim]Cl (5.00 g, 0.029 M) were dissolved separately in minimum amounts of ultrapure water. The [C₄mim]Cl solution was then slowly added to the stirred AgNO₃ solution. After stirring overnight the solution was filtered to remove the AgCl precipitate, the water removed and the IL dried under high vacuum at 70 °C overnight. The IL was then dissolved in 400 ml dry methanol, small amounts of activated charcoal and acidic alumina were added as seeds for the remaining AgCl, and the solution left overnight in a freezer. This solution was then filtered and the process repeated. The methanol was removed, and the ionic liquid was dried under high vacuum conditions. BZA (BDH Limited, 99.5%), ferrocene (Aldrich, 98%), tetrabutylammonium perchlorate (TBAP, Fluka, Puriss electrochemical grade, >99.99%) and MeCN (Fischer Scientific, dried and distilled, >99.99%) were used as received without further purification. Nitrogen gas (oxygen free) was purchased from BOC gases, Manchester, UK.

Instrumental

Electrochemical experiments were performed using a computer controlled μ -Autolab potentiostat (Eco-Chemie, The Netherlands). A conventional two-electrode system was employed for all experiments performed in RTILs, with a platinum electrode (10 or 2 μ m diameter) as the working electrode, and a 0.5 mm diameter silver wire as a quasi-reference electrode. The electrodes were housed in a glass cell 'T-cell' designed for investigating microsamples of ionic liquids under a controlled atmosphere.^[28,43] The working electrode was modified with a section of disposable micropipette tip to create a small cavity above the disk into which a drop (20–80 μ l) of ionic liquid (containing dissolved BZA) was placed. The RTIL solution was purged under vacuum (Edwards High Vacuum Pump, Model ES 50) for ca. 90 min, which served to remove trace atmospheric moisture and gases naturally present in the RTIL. Samples of different concentrations of BZA were systematically diluted in the T-cell with appropriate quantities of blank RTIL solvent. For experiments in MeCN, solutions were housed in a covered glass vial, with a three-electrode arrangement consisting of a 10 μ m diameter Pt working electrode, a silver wire reference electrode and Pt coil wire counter electrode. The platinum microdisk working electrodes were polished on self lapping pads (Kemet Ltd., UK) using alumina powder (Buehler, IL) of size 1.0 and 0.3 μ m. The electrode diameters were calibrated electrochemically by analysing the steady-state voltammetry of a 2 mM solution of ferrocene in MeCN containing 0.1 M TBAP, using a diffusion coefficient for ferrocene of $2.3 \times 10^{-5} \text{ cm}^2 \text{ s}^{-1}$ at 298 K.^[44] All experiments were performed inside a thermostated box (previously described by Evans *et al.*)^[45] which also functioned as a Faraday cage. The temperature was maintained at 298 (± 1.0) K.

Chronoamperometric experiments

Chronoamperometric transients were achieved using a sample time of 0.01 s (0.001 s in MeCN). After pre-equilibration for 20 s, the potential was stepped from a position of zero current to a chosen potential after the reductive peak, and the current was

measured for 10 s (0.5 s in MeCN). The software package Origin 7.0 (Microcal Software Inc.) was used to fit the experimental data. The equations proposed by Shoup and Szabo^[46] (below) were imported into the nonlinear curve fitting function, and the computer was instructed to perform 100 iterations on the data.

$$I = -4nFDcr_d f(\tau) \quad (3)$$

$$f(\tau) = 0.7854 + 0.8863\tau^{-\frac{1}{2}} + 0.2146 \exp(-0.7823\tau^{-\frac{1}{2}}) \quad (4)$$

$$\tau = \frac{4Dt}{r_d^2} \quad (5)$$

Here, n is the number of electrons transferred, F the Faraday constant, D the diffusion coefficient, c the initial concentration of parent species, r_d is the radius of the disk electrode and t the time. The equations used in this approximation are sufficient to give D and c within an error of 0.6%. The validity of this method in aprotic and RTIL solvents has also been described in a previous publication.^[47]

The value for the radius (previously calibrated) was fixed, and a value for the diffusion coefficient and the product of the number of electrons multiplied by concentration was obtained after optimization of the experimental data. It is noted that chronoamperometric transients in MeCN of larger duration than 0.5 s showed severe adsorption effects and could not be fitted to the Shoup and Szabo^[46] expression above.

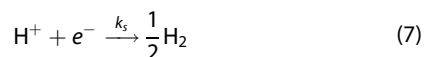
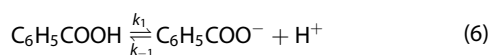
RESULTS AND DISCUSSION

The reduction of benzoic acid in various RTILs

The electrochemical reduction of BZA has been examined in six different ionic liquids, namely [C₂mim][NTf₂], [C₄mim][NTf₂], [C₄mpyr][NTf₂], [C₄mim][BF₄], [C₄mim][NO₃] and [C₄mim][PF₆], and MeCN. The six ionic liquids were chosen as suitable solvents for electrochemical experiments, since they showed wide, featureless baselines when purged in a vacuum for 90 min.

Cyclic voltammetry of BZA reduction in RTILs

Cyclic voltammetry for the reduction of BZA on a 10 μm diameter Pt electrode has been studied in six RTILs. Figure 2 shows typical voltammetry obtained at 100 and 1000 mV s⁻¹ for the reduction of a saturated solution of BZA in the RTIL [C₄mpyr][NTf₂]. Steady-state behaviour is observed at 100 mV s⁻¹, and the voltammetry becomes more transient in nature at higher scan rates (1000 mV s⁻¹). It is believed that the reduction peak at ca. -1.25 V (1000 mV s⁻¹) can be assigned to the one-electron CE reduction mechanism, where chemical dissociation of the acid precedes the electron-transfer step:^[31–34]



On the reverse sweep, particularly at higher scan rates, an oxidation peak is observed which suggests either a quasi-reversible process or the oxidation of a product of the reduction. It is believed that this peak is the oxidation of adsorbed hydrogen, consistent with a similar observation observed following the reduction of the ammonium ion in RTILs.^[30] There was no indication of a voltammetric peak for the oxidation of bulk

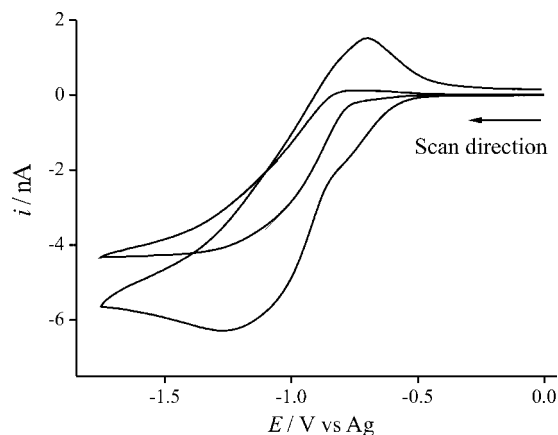


Figure 2. Typical cyclic voltammetry obtained for the reduction of a saturated solution of BZA in the RTIL [C₄mpyr][NTf₂] on a 10 μm diameter Pt electrode at 298 K, at scan rates of 100 and 1000 mV s⁻¹

hydrogen at all scan rates studied (10–4000 mV s⁻¹). There is also evidence of a 'pre-peak' prior to the main reduction peak, which is suggested to be due to reduction forming adsorbed hydrogen (H•(ads)). This peak, and the oxidation peak were both absent on gold electrodes, consistent with hydrogen being electrochemically inactive on gold.^[28] It should also be noted that the peaks observed are not thought to be due to the reduction of the BZA ring (which may be possible under some conditions); rather, protons are being reduced, leaving benzoate anions. A detailed mechanistic study in all six RTILs will be reported in a subsequent publication,^[48] but has not been included here since it is not the focus of this work. However, the focus will now turn to the measurement of chronoamperometric transients, both to calculate diffusion coefficients and solubilities, and to get an insight on the homogeneous kinetics of the CE reduction mechanism.

Potential step chronoamperometry

Potential-step chronoamperometry at a microdisk electrode is a commonly used electrochemical technique, and allows the independent calculation of diffusion coefficients and nc values (number of electrons multiplied by concentration) with a known microdisk radius. The approximation commonly used (see Eqns (3–5)) is described by Shoup and Szabo,^[46] and the validity of this approximation has recently been reported in aprotic and RTIL solvents,^[47] but only applies for a simple n electron-transfer process that is $A \pm ne^- \rightleftharpoons B$. In this work, it is believed that if the chemical step (Eqn (6), dissociation of BZA) is rate limiting, the experimental chronoamperometric transients will *not* fit the Shoup and Szabo approximation, and the limiting currents will *not* scale linearly with concentration.^[38]

To investigate this, chronoamperometric transients were obtained for the reduction of saturated solutions of BZA in six ionic liquids. The potential was stepped from a position of zero current (typically 0 V) to a potential after the reductive peak, and the current was measured for 10 s. In all cases, the theoretical fitting of the Shoup and Szabo expression^[46] to experimental transients was excellent ($\pm 0.7\%$) on a 10 μm diameter Pt electrode, implying that the C step is very fast, giving the response of a simple E step. The diffusion coefficients obtained were of the same order of magnitude ($10^{-11} \text{ m}^2 \text{ s}^{-1}$) as other solid species in RTILs (cf. $5.34 \times 10^{-11} \text{ m}^2 \text{ s}^{-1}$ for ferrocene in

Table 1. Chronoamperometric data obtained from theoretical fitting to the Shoup and Szabo^[46] expression for the electrochemical reduction of BZA in ionic liquids and MeCN at 298 K

RTIL/solvent	η at 293 K/cP	$D_{\text{BZA}} \times 10^{-11} \text{ m}^2 \text{ s}^{-1}$	Solubility/mM
[C ₂ mim][NTf ₂]	34 ^[40]	4.6 (± 0.3)	75 (± 2)
[C ₄ mim][NTf ₂]	52 ^[40]	3.2 (± 0.2)	78 (± 4)
[C ₄ mpyrr][NTf ₂]	89 ^[62]	2.7 (± 0.3)	74 (± 3)
[C ₄ mim][BF ₄]	112 ^[62]	1.8 (± 0.1)	220 (± 10)
[C ₄ mim][NO ₃]	266 ^[63]	0.26 (± 0.03)	2850 (± 200)
[C ₄ mim][PF ₆]	371 ^[63]	0.96 (± 0.05)	48 (± 3)
MeCN	0.34 ^[64]	170 (± 10)	850 ^{a[61]}

^a The solubility of BZA in MeCN (+0.1 M TBAP) could not be determined using chronoamperometric measurements in this work due to the formation of white crystals at high concentrations of BZA, possibly from a reaction of BZA with TBAP; D_{BZA} in MeCN was obtained from non-saturated solutions.

[C₂mim][NTf₂],^[49] and ca. 1–3 orders of magnitude less than that reported for BZA in water ($9.0 \times 10^{-10} \text{ m}^2 \text{ s}^{-1}$)^[50] and in near critical carbon dioxide ($1.3 \times 10^{-8} \text{ m}^2 \text{ s}^{-1}$).^[51]

The calculated diffusion coefficients obtained from this method are given in Table 1, and appear to decrease with increasing viscosity. The Stokes–Einstein relation^[52] given below (Eqn (8)) predicts a linear relationship between the diffusion coefficients (D) and the inverse of viscosity (η^{-1}) of simple diffusing species. This relationship is commonly followed in conventional molecular solvents, and also for large organic molecules,^[45] ferrocene^[49] and cobaltocenium hexafluorophosphate,^[49] but not for small molecules for example O₂,^[53] H₂,^[29] SO₂,^[54] NO₂^[55] and H₂S^[56] in RTILs.

$$D = \frac{kT}{6\eta\pi\alpha} \quad (8)$$

where k is the Boltzmann constant, T the temperature and α the hydrodynamic radius of the diffusing species. A plot of the diffusion coefficient of BZA *versus* the inverse of viscosity is shown in Fig. 3. An approximate linear relationship is observed (with [C₄mim][NO₃] an outlier at $0.26 \times 10^{-11} \text{ m}^2 \text{ s}^{-1}$), suggesting that Eqn (8) approximately holds for BZA in most RTILs.

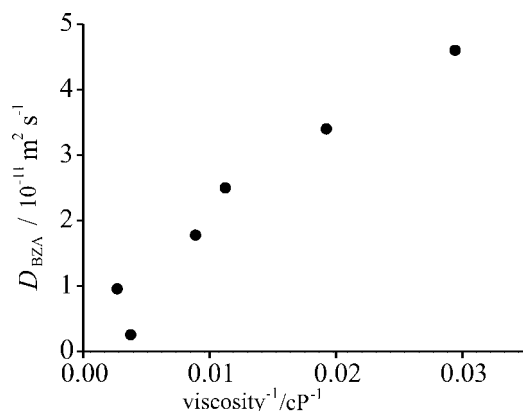


Figure 3. Plot of diffusion coefficient, D , (298 K) of BZA against the inverse of viscosity (293 K) for the six RTILs studied. D values were obtained from theoretical fitting of chronoamperometric transients to the Shoup and Szabo expression^[46]

Also included in Table 1 are the calculated solubilities of BZA in the six RTILs. The solubilities in the three [NTf₂]⁻ ionic liquids are almost exactly the same, and a slightly lower solubility is observed in [C₄mim][PF₆]. Interestingly, the solubility is higher in [C₄mim][BF₄], and significantly higher in [C₄mim][NO₃]. A strong interaction of protons with the nitrate anion *viz.* the formation of H[NO₃] or H[NO₃]₂⁻ (the hydrogen dinitrate ion) is conceivable. The existence of H[NO₃]₂⁻ is well known,^[57–60] and its formation in RTILs has been recently suggested.^[29] However, the chronoamperometric data implies the BZA is undissociated so we tentatively suggest a specific interaction of BZA with nitrate. In general, the solubilities in RTILs are relatively high compared to that reported in water (28 mM),^[61] but typically similar to than in MeCN (850 mM).^[61]

Reduction of different concentrations of BZA in [C₄mpyrr][NTf₂]

Since the chronoamperometric transients fit well in all RTILs for high concentrations (saturated solutions) of BZA, the transients were studied at different concentrations of BZA in the RTIL [C₄mpyrr][NTf₂]. It is expected that the steady-state (or chronoamperometric) limiting currents will not scale linearly with concentration for a CE process, except in the case of fast kinetics of the homogeneous step.^[37,38] Figure 4 shows typical experimental (solid lines) and fitted theoretical (dots) chronoamperometric transients for the reduction of BZA in [C₄mpyrr][NTf₂] at concentrations of 66, 55, 24 and 15 mM on a 10 μm diameter Pt electrode. Good fitting was observed at all concentrations, and the limiting current was found to increase linearly with the concentration of BZA (as shown in Fig. 6), suggesting that the homogeneous kinetics of Eqn (6) are indeed fast in RTILs.

Since it is also known that the kinetics of the homogeneous step can be studied by observing how the limiting current changes as a function of size,^[36–38] chronoamperometric transients were recorded at different concentrations of BZA on a smaller micro-electrode of diameter 2 μm . Figure 5 shows typical experimental (solid lines) and fitted theoretical (dots) chronoamperometric transients for the reduction of BZA in [C₄mpyrr][NTf₂] at concentrations of 41, 29, 17 and 5 mM. Consistent with the results obtained on the 10 μm diameter Pt electrode, good fitting was observed at all concentrations. A plot of the limiting current *versus* concentration of BZA on the 2 μm

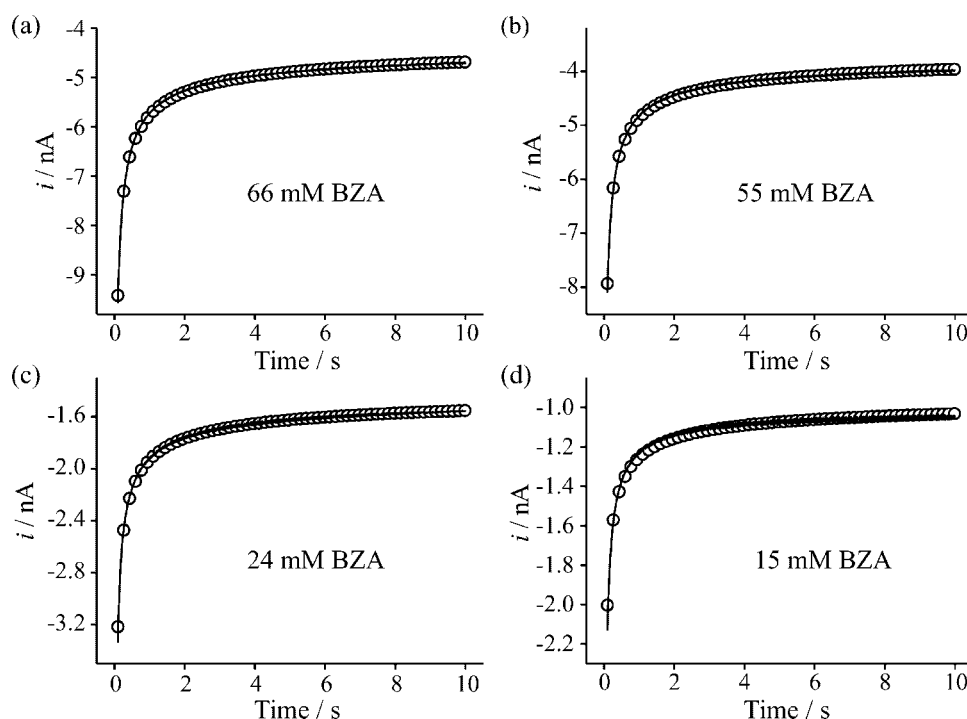


Figure 4. Experimental (solid lines) and fitted theoretical (dots) chronoamperometric transients for the reduction of BZA in $[\text{C}_4\text{mpyrr}][\text{NTf}_2]$ on a $10\ \mu\text{m}$ diameter Pt electrode at concentrations of (a) 66, (b) 55, (c) 24 and (d) 15 mM. The experimental data were theoretically fitted to the Shoup and Szabo^[46] expression following Eqns (3–5)

diameter Pt electrode is also included in Fig. 6. The gradient of the line on the $10\ \mu\text{m}$ diameter electrode is approximately five times that on the $2\ \mu\text{m}$ diameter electrode, suggesting that the limiting current is scaling with electrode size, supporting the suggestion

of fast kinetics of the homogeneous reaction given in Eqn (6), at least in $[\text{C}_4\text{mpyrr}][\text{NTf}_2]$.

In a previous paper, we used numerical simulation to find the limiting current, I_{lim} , for the CE process at a microdisk.^[36] The

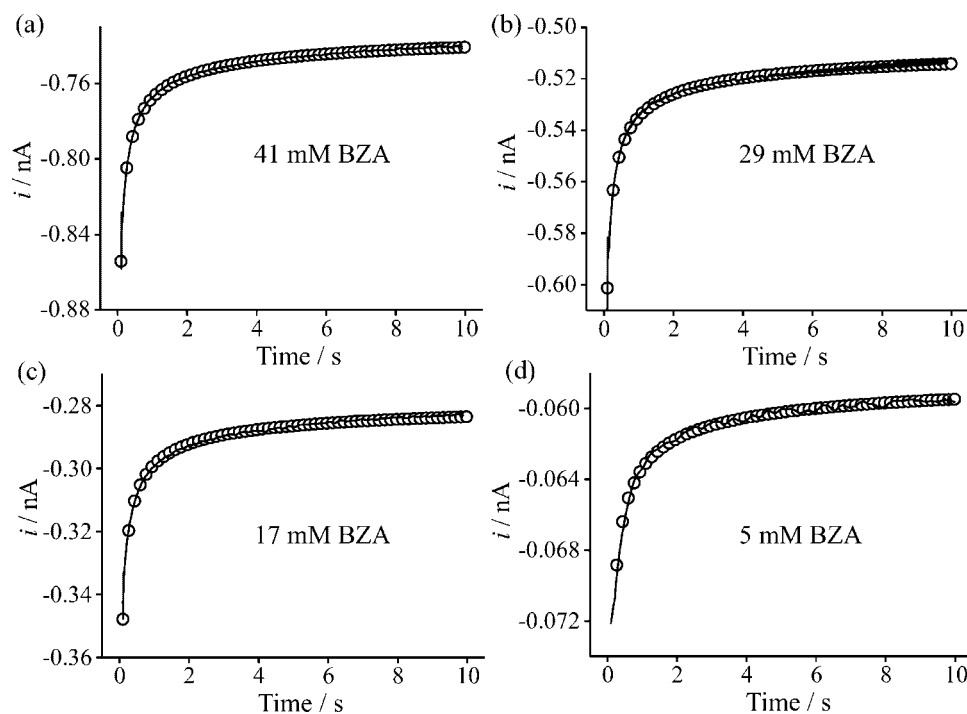


Figure 5. Experimental (solid lines) and fitted theoretical (dots) chronoamperometric transients for the reduction of BZA in $[\text{C}_4\text{mpyrr}][\text{NTf}_2]$ on a $2\ \mu\text{m}$ diameter Pt electrode at concentrations of (a) 41, (b) 29, (c) 17 and (d) 5 mM. The experimental data were theoretically fitted to the Shoup and Szabo^[46] expression following Eqns (3–5)

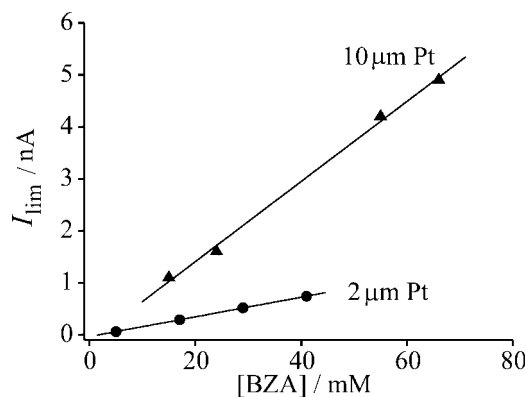


Figure 6. Plot of the chronoamperometric limiting current *versus* concentration for the reduction of BZA on a 10 and 2 μm diameter Pt electrode in $[\text{C}_4\text{mpyrr}][\text{NTf}_2]$ at 298 K

results were presented in terms of the value $I_{\text{lim}}/I_{\text{lim}}^*$, where I_{lim}^* is the limiting current that would be expected if there was no preceding homogeneous process. For a given value of disc radius, diffusion coefficient and concentration, the value $I_{\text{lim}}/I_{\text{lim}}^*$ is dependent only on the equilibrium constant and on the kinetics of the homogeneous process. When the homogeneous process has fast kinetics, or a large equilibrium constant, the value of $I_{\text{lim}}/I_{\text{lim}}^*$ is expected to be close to unity, whereas for slow kinetics its value is considerably less. The linear dependence of I_{lim} on the concentration of BZA suggests that we are working in a region where $I_{\text{lim}}/I_{\text{lim}}^*$ is close to unity. Figure 7 shows the region where $I_{\text{lim}}/I_{\text{lim}}^* > 0.95$, assuming a disc radius of 1.1 μm , a diffusion coefficient of $2.7 \times 10^{-11} \text{ m}^2 \text{ s}^{-1}$ for all species, and a concentration of 70 mM. This figure therefore gives an indication of the possible magnitudes of the rate constant k_1 for the deprotonation of BZA in $[\text{C}_4\text{mpyrr}][\text{NTf}_2]$.

Voltammetry for BZA reduction in MeCN

In order to compare the behaviour in ionic liquids to that obtained in a conventional aprotic solvent, the electrochemistry

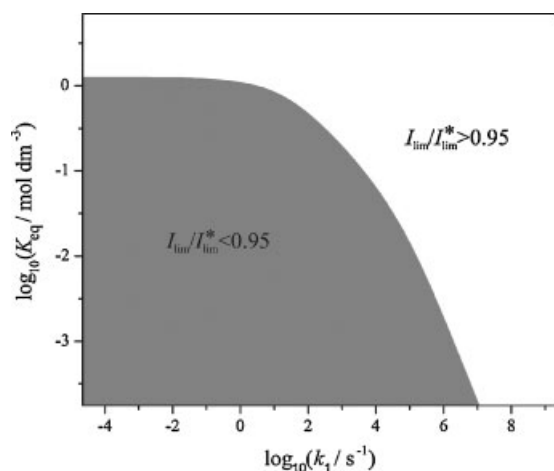


Figure 7. Map showing the values of K_{eq} and k_1 that would lead to the experimentally observed behaviour of the steady-state currents of BZA in $[\text{C}_4\text{mpyrr}][\text{NTf}_2]$. K_{eq} is defined as k_1/k_{-1}

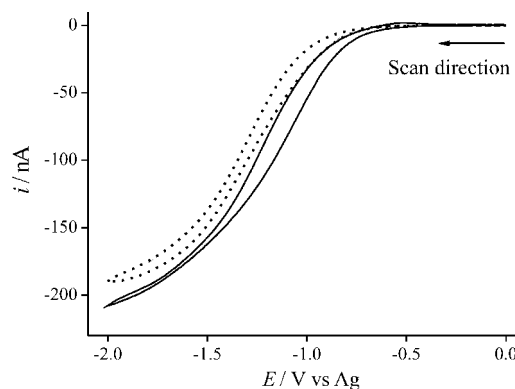


Figure 8. Typical cyclic voltammetry obtained for the reduction of 75 mM BZA in MeCN on a 10 μm diameter Pt electrode at 298 K, at scan rates of 100 (dots) and 1000 mV s^{-1}

of BZA reduction was also studied in MeCN. Figure 8 shows the reduction of 75 mM BZA in MeCN on a 10 μm Pt electrode at scan rates of 100 and 1000 mV s^{-1} . The voltammetry observed is steady-state in nature at all scan-rates studied (10–4000 mV s^{-1}) and no evidence of a 'pre-peak' or oxidation peak (seen in RTILs) was observed. In order to carry out the same kinetic analysis as in RTILs, chronoamperometric transients were recorded for the reduction of BZA at various different concentrations on a 10 μm diameter Pt electrode. In all cases, the theoretical fitting of experimental transients was very good, and a diffusion coefficient and concentration was obtained. The average diffusion coefficient obtained over the whole range of concentrations studied is given in Table 1 and is approximately two orders of magnitude faster than in RTILs, due to the lower viscosity of MeCN. The solubility could not be easily determined using chronoamperometric measurements due to the formation of crystals at saturated BZA concentrations, so a literature value is given.

Chronoamperometric transients were also recorded at different concentrations on a 2 μm diameter Pt electrode in MeCN, since the limiting current will only scale with concentration and electrode size if the homogeneous step is very fast.^[36–38] Figure 9 shows a plot of the chronoamperometric limiting current against the concentration of BZA in MeCN on a 10 and 2 μm diameter Pt electrode. On both electrodes, the limiting current increases linearly with scan rate, and the gradient of the line on the 10 μm

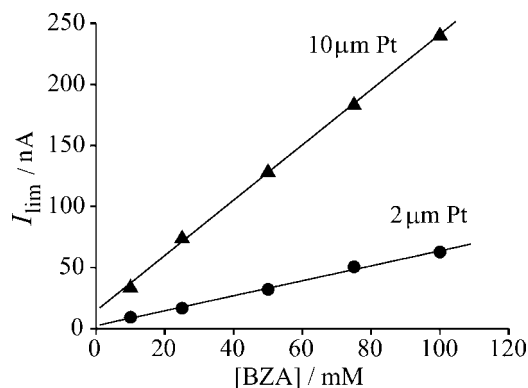


Figure 9. Plot of the chronoamperometric limiting current *versus* concentration for the reduction of BZA on a 10 and 2 μm diameter Pt electrode in MeCN at 298 K

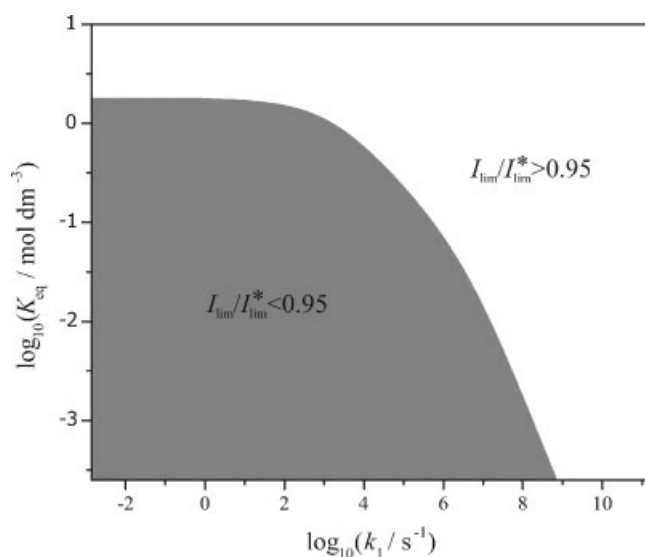


Figure 10. Map showing the values of K_{eq} and k_1 that would lead to the experimentally observed behaviour of the steady-state currents of BZA in MeCN. K_{eq} is defined as k_1/k_{-1}

is approximately five times that on the 2 μm . This suggests that, similar to the behaviour of BZA in RTILs, the C step given in Eqn (6) is also very fast in MeCN.

As in the last section, we can use the simulated results from our previous paper^[36] to interpret the steady state currents of BZA in MeCN. The linear dependence of I_{lim} on the concentration of BZA shows that we are working in a region where the value $I_{\text{lim}}/I_{\text{lim}}^*$ is near unity. Figure 10 shows the region where $I_{\text{lim}}/I_{\text{lim}}^* > 0.95$, assuming a disc radius of 1.1 μm , a diffusion coefficient of $1.7 \times 10^{-9} \text{ m}^2 \text{ s}^{-1}$ for all species, and a concentration of 100 mM. This figure therefore gives an indication of the possible magnitudes of the rate constant for the deprotonation of BZA in MeCN.

CONCLUSIONS

Due to the good agreement of theoretical chronoamperometric transients to experimental data, and the linear relationship observed between the limiting current and concentration, the electrochemical reduction of BZA appears to follow a CE mechanism in RTILs and MeCN, with a fast homogeneous chemical step preceding the electrochemical step. As a result, the diffusion coefficients and solubilities of BZA in several RTILs and MeCN have been calculated. A higher solubility of BZA was observed in $[\text{C}_4\text{mim}][\text{NO}_3]$, suggesting a strong interaction of the nitrate anion with the acid, otherwise the solubility in the RTILs are smaller/approximately similar to those in MeCN.

Acknowledgements

DSS thanks Schlumberger Cambridge Research and LA thanks the Department of Education and Learning in Northern Ireland and Merck GmbH for funding.

REFERENCES

- [1] K. N. Marsh, A. Deev, A. C.-T. Wu, E. Tran, A. Klamt, *Korean J. Chem. Eng.* **2002**, *19*, 357–362.
- [2] M. C. Buzzeo, R. G. Evans, R. G. Compton, *ChemPhysChem* **2004**, *5*, 1106–1120.
- [3] D. S. Silvester, R. G. Compton, *Z. Phys. Chem.* **2006**, *220*, 1247–1274.
- [4] F. Endres, S. Zein El Abedin, *PhysChemChemPhys* **2006**, *8*, 2101–2116.
- [5] M. J. Earle, K. R. Seddon, *Pure Appl. Chem.* **2000**, *72*, 1391–1398.
- [6] J. Dupont, C. Consorti, J. Spencer, *J. Braz. Chem. Soc.* **2000**, *11*, 337–344.
- [7] J. H. Davis, Jr. *Chem. Lett.* **2004**, *33*, 1072–1077.
- [8] J. H. Davis, Jr. *Synthesis of Task-Specific Ionic Liquids, in Ionic Liquids in Synthesis*, John Wiley & Sons, New York, USA, **2003**.
- [9] C. Jork, C. Kristen, D. Pieraccini, A. Stark, C. Chiappe, Y. A. Beste, W. J. Arlt, *Chem. Thermodyn.* **2005**, *37*, 537–558.
- [10] F. Endres, *ChemPhysChem* **2002**, *3*, 144–154.
- [11] P. Wasserscheid, W. Keim, *Angew. Chem. Int. Ed.* **2000**, *39*, 3772–3789.
- [12] A. J. Carmichael, M. J. Earle, J. D. Holbrey, P. B. McCormac, K. R. Seddon, *Org. Lett.* **1999**, *1*, 997–1000.
- [13] I. Meracz, T. Oh, *Tetrahedron Lett.* **2003**, *44*, 6465–6468.
- [14] H. Zhao, *Chem. Eng. Commun.* **2006**, *193*, 1660–1677.
- [15] M. C. Buzzeo, C. Hardacre, R. G. Compton, *Anal. Chem.* **2004**, *76*, 4583–4588.
- [16] N. Papageorgiou, Y. Athanassov, M. Armand, P. Bonhôte, H. Pettersson, A. Azam, M. Grätzel, *J. Electrochem. Soc.* **1996**, *143*, 3099–3108.
- [17] R. F. de Souza, J. C. Padilha, R. S. Goncalves, J. Dupont, *Electrochem. Commun.* **2005**, *5*, 728–731.
- [18] M. Ue, M. Takeda, A. Toriumi, A. Kominato, R. Hagiwara, Y. Ito, *J. Electrochem. Soc.* **2003**, *150*, A499–A502.
- [19] G. T. Sato, K. Masuda, K. Takagi, *Electrochim. Acta* **2004**, *49*, 3603–3611.
- [20] C. Thomazeau, H. Olivier-Bourbigou, L. Magna, S. Luts, B. Gilbert, *J. Am. Chem. Soc.* **2003**, *125*, 5264–5265.
- [21] D. R. MacFarlane, S. A. Forsyth, *ACS Symp. Ser.* **2003**, *856*, 264–276.
- [22] D. R. MacFarlane, J. M. Pringle, K. M. Johansson, S. A. Forsyth, M. Forsyth, *Chem. Commun.* **2006**, 1905–1917.
- [23] C. Hardacre, S. P. Katdare, D. Milroy, P. Nancarrow, D. W. Rooney, J. M. Thompson, *J. Catal.* **2004**, *227*, 44–52.
- [24] M. G. Del Pópolo, J. Kohanoff, R. M. Lynden-Bell, *J. Phys. Chem. B* **2006**, *110*, 8798–8803.
- [25] A. C. Cole, J. L. Jensen, I. Ntai, K. L. T. Tran, K. J. Weaver, D. C. Forbes, J. H. Davis, Jr. *J. Am. Chem. Soc.* **2002**, *124*, 5962–5963.
- [26] Y. Du, F. Tian, *Synth. Commun.* **2005**, *35*, 2703–2708.
- [27] M. A. B. H. Susan, A. Noda, S. Mitsushima, M. Watanabe, *Chem. Commun.* **2003**, 938–939.
- [28] D. S. Silvester, L. Aldous, C. Hardacre, R. G. Compton, *J. Phys. Chem. B* **2007**, *111*, 5000–5007.
- [29] D. S. Silvester, K. R. Ward, L. Aldous, C. Hardacre, R. G. Compton, *J. Electroanal. Chem.* **2008**, *618*, 53–60.
- [30] X. Ji, D. S. Silvester, L. Aldous, C. Hardacre, R. G. Compton, *J. Phys. Chem. C* **2007**, *111*, 9562–9572.
- [31] W. C. Barrette, Jr. D. T. Sawyer, *Anal. Chem.* **1984**, *56*, 653–657.
- [32] S. E. Treimer, D. H. Evans, *J. Electroanal. Chem.* **1998**, *455*(1–2), 19–27.
- [33] W. Vielstich, D. Jahn, *Z. Elektrochem.* **1960**, *64*, 43–44.
- [34] W. J. Albery, R. P. Bell, *Proc. Chem. Soc.* **1963**, *6*, 169–170.
- [35] R. G. Barradas, O. Kutowy, D. W. Shoosmith, *Electrochim. Acta.* **1974**, *19*, 49–56.
- [36] I. Streeter, R. G. Compton, *J. Electroanal. Chem.* **2008**, *615*, 154–158.
- [37] S. Daniele, I. Lavagnini, M. Baldo, F. Magno, *Anal. Chem.* **1998**, *70*, 285–294.
- [38] S. Daniele, M. Baldo, C. Bragato, M. Abdelsalam, G. Denuault, *Anal. Chem.* **2002**, *74*, 3290–3296.
- [39] M. Fleischmann, F. Lasserre, J. Robinson, D. Swan, *J. Electroanal. Chem.* **1984**, *177*(1–2), 97–114.
- [40] P. Bonhôte, A.-P. Dias, N. Papageorgiou, K. Kalyanasundaram, M. Grätzel, *Inorg. Chem.* **1996**, *35*, 1168–1178.
- [41] D. R. MacFarlane, P. Meakin, J. Sun, N. Amini, M. Forsyth, *J. Phys. Chem. B* **1999**, *103*, 4164–4170.
- [42] L. Cammarata, S. G. Kazarian, P. A. Salter, T. Welton, *PhysChemChemPhys* **2001**, *3*, 5192–5200.
- [43] U. Schröder, J. D. Wadhawan, R. G. Compton, F. Marken, P. A. Z. Suarez, C. S. Consorti, R. F. de Souza, J. Dupont, *New J. Chem.* **2000**, *24*, 1009–1015.
- [44] M. Sharp, *Electrochim. Acta* **1983**, *28*, 301–308.

- [45] R. G. Evans, O. V. Klymenko, P. D. Price, S. G. Davies, C. Hardacre, R. G. Compton, *ChemPhysChem* **2005**, *6*, 526–533.
- [46] D. Shoup, A. Szabo, *J. Electroanal. Chem. Interfacial Electrochem.* **1982**, *140*, 237–245.
- [47] C. A. Paddon, D. S. Silvester, F. L. Bhatti, T. J. Donohoe, R. G. Compton, *Electroanalysis* **2007**, *19*, 11–22.
- [48] D. S. Silvester, W. He, L. Aldous, C. Hardacre, R. G. Compton, The Electrochemical Reduction of Benzoic Acid and Substituted Benzoic Acid in Some Room Temperature Ionic Liquids. *J. Phys. Chem. C* **2008**, in press.
- [49] E. I. Rogers, D. S. Silvester, D. L. Poole, L. Aldous, C. Hardacre, R. G. Compton, *J. Phys. Chem. C* **2008**, *112*, 2729–2735.
- [50] T. Hotta, S. Nii, T. Yajima, F. Kawaizumi, *Chem. Eng. Technol.* **2007**, *30*, 208–213.
- [51] O. J. Catchpole, M. B. King, *Ind. Eng. Chem. Res.* **1994**, *33*, 1828–1837.
- [52] R. G. Compton, C. E. Banks, *Understanding Voltammetry*, World Scientific, Singapore, **2007**.
- [53] R. G. Evans, O. V. Klymenko, S. A. Saddoughi, C. Hardacre, R. G. Compton, *J. Phys. Chem. B* **2004**, *108*, 7878–7886.
- [54] L. E. Barrosse-Antle, D. S. Silvester, L. Aldous, C. Hardacre, R. G. Compton, *J. Phys. Chem. B* **2008**, *112*, 3398–3404.
- [55] T. L. Broder, D. S. Silvester, L. Aldous, C. Hardacre, R. G. Compton, *J. Phys. Chem. B* **2007**, *111*, 7778–7785.
- [56] A. M. O'Mahony, D. S. Silvester, L. Aldous, C. Hardacre, R. G. Compton, *J. Phys. Chem. C* **2008**, *112*, 7725–7730.
- [57] G. Dobinson, R. Mason, D. Russell, *Chem. Commun.* **1967**, 62–63.
- [58] B. Faithful, R. Gillard, D. Tuck, R. Ugo, *J. Chem. Soc. A* **1966**, 1185–1188.
- [59] A. Sengul, *Turk. J. Chem.* **2005**, *29*, 571–578.
- [60] L. Siew, B. Sundheim, *J. Phys. Chem.* **1969**, *73*, 4135–4141.
- [61] M. Chantooni, Jr I. Kolthoff, *J. Phys. Chem.* **1974**, *78*, 839–846.
- [62] O. O. Okoturo, T. J. VanderNoot, *J. Electroanal. Chem.* **2004**, *568*, 167–181.
- [63] K. R. Seddon, A. Stark, M.-J. Torres, *ACS Symp. Ser.* **2002**, *819*, 34–39.
- [64] Ed.: D. R. Lide, *Handbook of Chemistry and Physics* (76th edn), CRC Press, Boca Raton, USA. **1996**.